

Hardness in Drinking-water

Background document for development of
WHO Guidelines for Drinking-water Quality

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Background document for development of WHO *Guidelines for Drinking-water Quality*

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Preface

One of the primary goals of the World Health Organization (WHO) and its Member States is that “all people, whatever their stage of development and their social and economic conditions, have the right to have access to an adequate supply of safe drinking water”. A major WHO function to achieve such goals is the responsibility “to propose ... regulations, and to make recommendations with respect to international health matters”

The first WHO document dealing specifically with public drinking-water quality was published in 1958 as *International Standards for Drinking-water*. It was subsequently revised in 1963 and in 1971 under the same title. In 1984–1985, the first edition of the *WHO Guidelines for Drinking-water Quality* (GDWQ) was published in three volumes: Volume 1, Recommendations; Volume 2, Health criteria and other supporting information; and Volume 3, Surveillance and control of community supplies. Second editions of these volumes were published in 1993, 1996 and 1997, respectively. Addenda to Volumes 1 and 2 of the second edition were published in 1998, addressing selected chemicals. An addendum on microbiological aspects reviewing selected microorganisms was published in 2002. The third edition of the GDWQ was published in 2004, the first addendum to the third edition was published in 2006 and the second addendum to the third edition was published in 2008. The fourth edition will be published in 2011.

The GDWQ are subject to a rolling revision process. Through this process, microbial, chemical and radiological aspects of drinking-water are subject to periodic review, and documentation related to aspects of protection and control of public drinking-water quality is accordingly prepared and updated.

Since the first edition of the GDWQ, WHO has published information on health criteria and other supporting information to the GDWQ, describing the approaches used in deriving guideline values and presenting critical reviews and evaluations of the effects on human health of the substances or contaminants of potential health concern in drinking-water. In the first and second editions, these constituted Volume 2 of the GDWQ. Since publication of the third edition, they comprise a series of free-standing monographs, including this one.

For each chemical contaminant or substance considered, a lead institution prepared a background document evaluating the risks for human health from exposure to the particular chemical in drinking-water. Institutions from Canada, Japan, the United Kingdom and the United States of America (USA) prepared the documents for the fourth edition.

Under the oversight of a group of coordinators, each of whom was responsible for a group of chemicals considered in the GDWQ, the draft health criteria documents were submitted to a number of scientific institutions and selected experts for peer review. Comments were taken into consideration by the coordinators and authors. The draft documents were also released to the public domain for comment and submitted for final evaluation by expert meetings.

During the preparation of background documents and at expert meetings, careful consideration was given to information available in previous risk assessments carried out by the International Programme on Chemical Safety, in its Environmental Health Criteria monographs and Concise International Chemical Assessment Documents, the International Agency for Research on Cancer, the Joint FAO/WHO Meeting on Pesticide Residues and the Joint FAO/WHO Expert Committee on Food Additives (which evaluates contaminants such as lead, cadmium, nitrate and nitrite, in addition to food additives).

Further up-to-date information on the GDWQ and the process of their development is available on the WHO Internet site and in the current edition of the GDWQ.

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The work of the following working group coordinators was crucial in the development of this document and others contributing to the fourth edition:

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The draft text was discussed at the Expert Consultation for the fourth edition of the GDWQ, held in December 2010. The final version of the document takes into consideration comments from both peer reviewers and the public. The input of those who provided comments and of participants at the meeting is gratefully acknowledged.

The WHO coordinators were Mr R. Bos and Mr B. Gordon, WHO Headquarters. Ms C. Vickers provided a liaison with the International Programme on Chemical Safety, WHO Headquarters. Mr M. Zaim, Public Health and the Environment Programme, WHO Headquarters, provided input on pesticides added to drinking-water for public health purposes.

Ms P. Ward provided invaluable administrative support throughout the review and publication process. Ms M. Sheffer of Ottawa, Canada, was responsible for the scientific editing of the document.

Many individuals from various countries contributed to the development of the GDWQ. The efforts of all who contributed to the preparation of this document and in particular those who provided peer or public domain review comments are greatly appreciated.

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Much of this background document was extracted from the book *Calcium and magnesium in drinking-water: public health significance* (WHO, 2009). The interested reader should consult that publication for additional information and for primary references.

1. GENERAL DESCRIPTION

1.1 Identity

Water hardness is the traditional measure of the capacity of water to react with soap, hard water requiring considerably more soap to produce a lather. Hard water often produces a noticeable deposit of precipitate (e.g. insoluble metals, soaps or salts) in containers, including “bathtub ring”. It is not caused by a single substance but by a variety of dissolved polyvalent metallic ions, predominantly calcium and magnesium cations, although other cations (e.g. aluminium, barium, iron, manganese, strontium and zinc) also contribute. Hardness is most commonly expressed as milligrams of calcium carbonate equivalent per litre. Water containing calcium carbonate at concentrations below 60 mg/l is generally considered as soft; 60–120 mg/l, moderately hard; 120–180 mg/l, hard; and more than 180 mg/l, very hard (McGowan, 2000). Although hardness is caused by cations, it may also be discussed in terms of carbonate (temporary) and non-carbonate (permanent) hardness.

1.2 Sources

The principal natural sources of hardness in water are dissolved polyvalent metallic ions from sedimentary rocks, seepage and runoff from soils. Calcium and magnesium, the two principal ions, are present in many sedimentary rocks, the most common being limestone and chalk. They are also common essential mineral constituents of food. As mentioned above, a minor contribution to the total hardness of water is also made by other polyvalent ions, such as aluminium, barium, iron, manganese, strontium and zinc.

1.3 Organoleptic properties

The taste threshold for the calcium ion is in the range 100–300 mg/l, depending on the associated anion, but higher concentrations are acceptable to consumers (see also section 4.1).

2. ENVIRONMENTAL LEVELS AND HUMAN EXPOSURE

2.1 Water

Small water supplies using groundwater often encounter significant levels of hardness, but some larger surface water supplies also have the same issue. Calcium concentrations up to and exceeding 100 mg/l are common in natural sources of water, particularly groundwater. Magnesium is present in natural groundwater usually at lower concentrations (from negligible to about 50 mg/l and rarely above 100 mg/l), so calcium-based hardness usually predominates (National Research Council, 1977).

Estimated daily intakes of magnesium from water of about 2.3 mg and 52.1 mg in soft-water and hard-water areas, respectively, have been reported, based on adults drinking 2 litres of water per day (Neri et al., 1985).

2.2 Food

Food is the principal dietary source of intake of both calcium and magnesium. Dairy products are the richest sources of dietary calcium, contributing over 50% of the total calcium in many diets. Some plant foods, including legumes, green leafy vegetables and broccoli, can also contribute to dietary calcium, but the content is lower than in dairy products, and the bioavailability of calcium (and magnesium) in plant foods can be low if the concentration of oxalate or phytate is high. Dietary sources of magnesium are more varied; dairy products, vegetables, grains, fruits and nuts are important contributors.

Typical recommended dietary intakes are about 1000 mg of calcium per day and 200–400 mg of magnesium per day. Moves to reduce the intake of dairy products because of the fat content will lead to a lowering of calcium and magnesium intakes in some population groups. Populations that use only a very small amount of dairy products would also have a lower intake of calcium.

2.3 Estimated total exposure and relative contribution of drinking-water

The typical dietary contribution of calcium and magnesium is over 80% of the total daily intake. Of this, approximately 30% of calcium and 35% of magnesium will be absorbed. The bioavailabilities of calcium and magnesium from milk and water are on the order of 50% (Ong, Grandjean & Heaney, 2009). For calcium and magnesium, the typical contribution from water is 5–20% (WHO, 1973; National Research Council, 1977; Neri & Johansen, 1978).

Because of dietary habits in most countries, many people fail to obtain the recommended intakes of one or both of these nutrients from their diets. While the concentrations of calcium and magnesium in drinking-water vary markedly from one supply to another, mineral-rich drinking-waters may provide substantial contributions to total intakes of these nutrients in some populations or population subgroups. Water treatment processes can affect mineral concentrations and, hence, the total intake of calcium and magnesium for some individuals.

3. EFFECTS ON HUMANS

Both calcium and magnesium are essential minerals and beneficial to human health in several respects. Inadequate intake of either nutrient can result in adverse health consequences. Recommended daily intakes of each element have been set at national and international levels. Individuals vary considerably in their needs for and consumption of these elements.

3.1 Calcium

3.1.1 Inadequate intake

Inadequate intakes of calcium have been associated with increased risks of osteoporosis, nephrolithiasis (kidney stones), colorectal cancer, hypertension and stroke, coronary artery disease, insulin resistance and obesity. Most of these disorders have treatments, but not cures. Owing to a lack of compelling evidence for the role of calcium as a contributory element in relation to these diseases, estimates of calcium requirement have been made on the basis of bone health outcomes, with the goal of optimizing bone mineral density.

Calcium is unique among nutrients, in that the body's reserve is also functional; increasing bone mass is linearly related to reduction in fracture risk. The total body stores are on the order of 1200 g, with about 99% in bones and teeth. A large body of primary evidence from randomized controlled trials shows that increasing calcium intake, especially in those who have had habitually low calcium intakes, increases bone mass during growth and reduces bone loss and fracture risk late in life. Osteoporosis is one of the most prevalent of age-related diseases. Calcium and vitamin D are jointly beneficial in increasing bone mass.

Epidemiological evidence indicates that dietary calcium reduces the incidence of kidney stones. In contrast, the results of a large randomized trial suggest an increased risk of kidney stones associated with calcium supplements, possibly because the calcium was ingested as a bolus and not with food or the supplements were taken by those who exceeded the upper intake level of 2500 mg/day.

Although hypertension is multifactorial in origin, adequate calcium intake has been associated with lowered risk of elevated blood pressure in some, but not all, studies. A clear mechanism has not been identified, although electrolytes probably play a role. Dairy products, more than calcium per se, have been associated with reduced blood pressure in randomized prospective studies and with reduced risk of stroke in prospective studies.

Those individuals who avoid dairy products or lack access to them throughout life may be at increased risk of calcium deficiency. Formula-fed infants will not normally be at risk from deficient or excess amounts of calcium, as even extremely low or high calcium concentrations in water would not lead to absorption of non-physiological amounts of calcium from infant formula reconstituted with the water. If, however, other food sources are used that do not provide the calcium content of full-strength formula, then water may represent an important source of the mineral for the infants.

3.1.2 Excess intake

To a great extent, individuals are protected from excess intakes of calcium by a tightly regulated intestinal absorption and elimination mechanism through the action of 1,25-dihydroxyvitamin D, the hormonally active form of vitamin D. When calcium is absorbed in excess of need, the excess is excreted by the kidney in healthy people who do not have renal impairment. Concern for excess calcium intake is directed primarily to those who are prone to milk alkali syndrome (the simultaneous presence

of hypercalcaemia, metabolic alkalosis and renal insufficiency) and hypercalcaemia. Although calcium can interact with iron, zinc, magnesium and phosphorus within the intestine, thereby reducing the absorption of these minerals, available data do not suggest that these minerals are depleted when humans consume diets containing calcium above the recommended levels. For example, even though high intakes of calcium can exert acute effects on iron absorption, there is no evidence of reduced iron status or iron stores with long-term calcium supplementation.

3.2 Magnesium

3.2.1 Inadequate intake

Magnesium is the fourth most abundant cation in the body and the second most abundant cation in intracellular fluid. It is a cofactor for some 350 cellular enzymes, many of which are involved in energy metabolism. It is also involved in protein and nucleic acid synthesis and is needed for normal vascular tone and insulin sensitivity. Total body stores are on the order of 25 g, with about 60% in bone. Total body burden is difficult to quantify, because only a small portion is in blood or fluids, and it can be variable. Low magnesium levels are associated with endothelial dysfunction, increased vascular reactions, elevated circulating levels of C-reactive protein (a proinflammatory marker that is a risk factor for coronary heart disease) and decreased insulin sensitivity. Low magnesium status has been implicated in hypertension, coronary heart disease, type 2 diabetes mellitus and metabolic syndrome.

Magnesium deficiency has been implicated in the pathogenesis of hypertension, with some epidemiological and experimental studies demonstrating a negative correlation between blood pressure and serum magnesium levels. However, data from clinical studies have been less convincing. Cardiac arrhythmias of ventricular and atrial origin have been reported in patients with hypomagnesaemia and in postmenopausal women in controlled diet studies. Indeed, a serious cardiac arrhythmia, Torsade de Pointes, is treated with intravenous magnesium therapy.

Pre-eclampsia (defined as hypertension after 20 weeks of gestation) with proteinuria has been treated with magnesium salts for many decades. A recent trial (Altman et al., 2002) using magnesium sulfate showed a 50% decreased risk of eclampsia.

Animal studies have documented an inverse (protective) relationship between magnesium intake and the rate or incidence of atherosclerosis. In humans, there is evidence for an inverse (protective) relationship between magnesium and coronary heart disease mortality. Three cross-sectional studies have now documented an inverse relationship between the concentration of C-reactive protein and magnesium intake or serum magnesium concentration, suggesting that magnesium may have an anti-inflammatory effect.

Several studies have documented the importance of magnesium in type 2 diabetes mellitus. Two recent studies have documented an inverse (protective) relationship between magnesium intake and risk of developing type 2 diabetes mellitus. Oral magnesium supplementation improves insulin sensitivity and metabolic control in type 2 diabetes mellitus.

Alcoholism and intestinal malabsorption are conditions associated with magnesium deficiency. Some drugs, such as certain diuretics, some antibiotics and some chemotherapy treatments, increase the loss of magnesium through the kidney; therefore, those patients should have magnesium supplementation as part of their therapy.

3.2.2 Excess intake

The major cause of hypermagnesaemia is renal insufficiency associated with a significantly decreased ability to excrete magnesium. Increased intake of magnesium salts may cause a temporary adaptable change in bowel habits (diarrhoea), but seldom causes hypermagnesaemia in persons with normal kidney function. Drinking-water in which both magnesium and sulfate are present at high concentrations (above approximately 250 mg/l each) can have a laxative effect, although data suggest that consumers adapt to these levels as exposures continue. Laxative effects have also been associated with excess intake of magnesium taken in the form of supplements, but not with magnesium in diet.

3.3 Epidemiological studies

A large number of studies have investigated the potential beneficial health effects of drinking-water hardness. Most of these have been ecological epidemiological studies and have reported an inverse relationship between water hardness and cardiovascular mortality. Inherent weaknesses in the ecological epidemiological study design limit the conclusions that can be drawn from these studies.

Several identified case-control and cohort studies show a negative association (i.e. protective effect) between cardiovascular mortality and drinking-water magnesium. Although this association does not necessarily demonstrate causality, it is consistent with the well-known effects of magnesium on cardiovascular function. There was no evidence of an association between total water hardness or calcium and acute myocardial infarction or deaths from cardiovascular disease (acute myocardial infarction, stroke and hypertension). There does not appear to be an association between drinking-water magnesium and acute myocardial infarction. A recent large study from the Netherlands (Leurs et al., 2010) found no overall association between calcium, magnesium or total hardness and ischaemic heart disease or stroke mortality. However, there was a reported significant inverse (beneficial) association with water magnesium for men in the highest exposure group, and the opposite effect was observed for women. Thus, further study is needed.

Case-control and cohort studies are more useful than ecological epidemiological studies for investigating cause-and-effect relationships. Seven case-control studies and two cohort studies of acceptable quality investigating the relationship between calcium or magnesium and cardiovascular disease or mortality were identified in the literature. Of the case-control studies, one addressed the association between calcium and acute myocardial infarction and three the association between calcium and death from cardiovascular disease. None found a positive or inverse correlation between calcium and either morbidity or mortality. Two examined the relationship between magnesium and acute myocardial infarction, finding no association. Five examined the relationship between magnesium and cardiovascular mortality; while some failed

to yield statistically significant results, collectively they showed similar trends of reduced cardiovascular mortality as magnesium concentrations in water increased. Statistically significant benefits (where observed) generally occurred at magnesium concentrations of about 10 mg/l and greater. The cohort studies examined the relationship between water hardness (rather than calcium or magnesium content) and cardiovascular disease or mortality and found no association.

3.4 Other health effects

Exposure to hard water has been suggested to be a risk factor that could exacerbate eczema. The environment plays an important part in the etiology of atopic eczema, but specific causes are unknown. Numerous factors have been associated with eczema flare-up, including dust, nylon, shampoo, sweating, swimming and wool (Langan, 2009). A suggested explanation relative to hard water is that increased soap usage in hard water results in metal or soap salt residues on the skin (or on clothes) that are not easily rinsed off and that lead to contact irritation (Thomas & Sach, 2000). There are reports of a relationship between both 1-year and lifetime prevalence of atopic eczema and water hardness among primary-school children. Eczema prevalence trends in the secondary-school population were not significant (McNally et al., 1998). Additional studies are under way.

4. OTHER CONSIDERATIONS

4.1 Taste acceptability

Dissolved minerals contribute to the taste of drinking-water to varying degrees. Acceptability of water will usually depend on the individual user's taste and familiarity. Demineralized water tends to have a flat taste, and producers of demineralized bottled or packaged water often add some minerals for taste. Some bottled mineral waters have exceptionally high mineral concentrations and appeal to some consumers, but would not be considered acceptable for most public drinking-water supplies. Concentrations of calcium and magnesium and other dissolved solids in water that are detectable by consumers are manageable by treatment or blending in public drinking-waters.

4.2 Corrosion and scaling

Depending upon interactions with other factors, such as pH and alkalinity, hard water can cause increased soap consumption and scale deposition in the water distribution system, as well as in heated water applications where insoluble metal carbonates are formed, coating surfaces and reducing the efficiency of heat exchangers. Excessively hard water can also have corrosion tendencies. Soft water that is not stabilized has a great tendency to cause corrosion of metal surfaces and pipes, resulting in the presence of certain heavy metals, such as cadmium, copper, lead and zinc, in drinking-water (National Research Council, 1977). Corrosion can be associated with health risks (from leachates such as lead, copper and other metals) and reduced lifespan of the distribution network and appliances (e.g. water heaters) using water. Soft or softened waters do have the benefit of minimal scaling and therefore allow more efficient heat transfer in exchangers and probably longer life of hot water heaters.

4.3 Conditioning of water

When conditioning is done, the target is normally to achieve bicarbonate equilibrium and suitable pH and alkalinity. Other ions, such as sulfate, nitrate and chloride, can be implicated in corrosivity.

Central water softening treatment usually involves lime (hydrated calcium oxide) or lime soda (lime plus sodium carbonate) softening and is commonly practised. These chemicals increase the precipitation of the calcium and magnesium carbonate, reducing the calcium hardness of the treated water. These waters are balanced to minimize post-precipitation of lime and should be stabilized as needed to control corrosivity.

There are significant differences between naturally soft water, including rainwater or soft deionized water, and water that has been softened by cation exchange, wherein the divalent cations (Ca^{2+} , Mg^{2+} , etc.) have been replaced by sodium. Cation exchange-softened water, although containing substantial amounts of sodium and chloride, is not necessarily corrosive. Naturally soft waters also require similar stabilization and corrosion reduction treatment prior to distribution.

4.4 Desalination

Desalination of seawater and brackish water converts water with a high dissolved solids content to water with a very low dissolved solids content. Treatment processes that desalinate water by nanofiltration, which is fairly selective for removal of divalent ions, reverse osmosis membranes or thermal desalination will deplete the mineral content and increase corrosivity. These demineralized waters are highly aggressive and must be stabilized prior to distribution by addition of lime and alkalinity and possibly chelators, such as phosphate. Some desalinated waters are back-blended with source waters to increase the mineralization somewhat and reduce corrosivity (Cotruvo et al., 2010).

Stabilization practices should ensure that the overall process does not significantly reduce the total intake of nutrients such as calcium and magnesium below recommended values. Based on local circumstances, water suppliers and public health authorities may wish to further modify final drinking-water composition in light of overall mineral nutrition.

4.5 Reuse water

Indirect wastewater reuse involves the extraction of water from sources that have inputs from wastewater discharges. Planned indirect potable reuse of wastewater, where wastewater discharges are sited close to drinking-water extraction points, is a growing source of drinking-water in some localities. For such planned reuse, enhanced treatment steps are usually employed.

The total dissolved solids content of domestic wastewater is greater than that of the original drinking-water. In some settings, the wastewater is treated by membrane technologies to reduce the levels of total dissolved solids, as well as for purification.

If groundwater recharge or groundwater storage is part of the process, additional stabilization may be needed after withdrawal of the water and prior to its distribution.

Treatment and stabilization practices should ensure that the overall process does not significantly reduce total intake of nutrients such as calcium and magnesium below recommended values. Based on local circumstances, water suppliers and public health authorities may wish to further modify final drinking-water composition in light of overall mineral nutrition.

4.6 Packaged water

Packaged waters can be spring or mineral waters or bottled tap waters. Because of extreme variation in the mineral composition of marketed bottled waters, with levels of total dissolved solids ranging from almost zero to several thousand milligrams per litre and with a similar variation in concentrations of essential elements, the public should have access to information on the mineral composition of bottled or packaged water.

4.7 Naturally soft water

Naturally soft water can have aggressive properties towards the piping material through which it is distributed. To avoid the corrosion of piping materials, the water is normally conditioned or stabilized. Frequently, this involves increasing the alkalinity and/or adding corrosion-inhibiting substances (e.g. phosphates) in some form. The choice for the most appropriate conditioning technology will depend on local circumstances (e.g. water quality issues, piping materials, corrosion). Based on local circumstances, water suppliers and public health authorities may wish to further modify drinking-water composition in light of overall mineral nutrition.

4.8 Collected rainwater

Rainwater collection refers to collection at the household or local community level for local use. Rainwater is soft and usually slightly acidic. If it is distributed through a piped system, the same considerations as for naturally soft water apply. In some settings, marble chips (calcium carbonate) are added to rainwater storage tanks. This will contribute to calcium intake and corrosion prevention.

4.9 Household water treatment

Point-of-entry ion exchange (water softener) devices are used in some households to remove hardness (calcium, magnesium) and iron from water. Each divalent ion (e.g. Ca^{2+} or Mg^{2+}) in the water is replaced by two sodium ions. Softening will have several aesthetically beneficial effects inside the home, such as reducing scaling in pipes, fixtures and water heaters and improving laundry and washing characteristics, but it also increases the sodium (and chloride) content of the drinking-water. Consumption of calcium and magnesium in drinking-water will, of course, be lower unless the water that is consumed is not softened or is remineralized.

Point-of-use reverse osmosis and distilling devices remove virtually all the minerals from the input water, and they can remove several types of potential trace-level contaminants that may be present, as well as removing nutrient minerals. While this

water need not be conditioned if materials not subject to corrosion are used after the treatment, the resultant drinking-water is devoid of minerals. So, use of these devices may result in the reduction of the overall intake of nutrient minerals by the consumers in the households.

Users of these devices should be made aware of the changes in mineral composition that arise and the possible consequences for total nutrient intake. For example, those who sell or install these devices may be encouraged to bring to the attention of the users of these devices the possibility of reduced intake of minerals.

One approach to ensure that the water used for drinking and cooking is not demineralized is to soften only the hot water line at the entry to the hot water heater, which provides several benefits and also reduces costs. Additionally, the manufacturers of these water softeners may provide a suitable bypass of a portion of this water to maintain some level of these minerals in the water actually consumed (e.g. to a kitchen tap) or develop and add an appropriate remineralizing unit in the water line prior to the point of consumption.

5. CONCLUSIONS

Natural and treated waters have a wide range of mineral content, from very low levels in rainwaters and naturally soft and softened water to moderate and very high levels in naturally hard waters and waters with high total dissolved solids content. Bottled and packaged waters can be naturally mineralized or naturally soft or demineralized. Thus, the mineral consumption from drinking-water and cooking water will vary widely, depending upon location, treatment and water source.

The degree of hardness of drinking-water is important for aesthetic acceptability by consumers and for economic and operational considerations. Many hard waters are softened for those reasons using several applicable technologies, and the mineral composition will be significantly affected. The choice for the most appropriate conditioning technology will depend on local circumstances (e.g. water quality issues, piping materials, corrosion). Some softening is provided at the central treatment plant, and some is provided in individual homes as a consumer preference. Modification of calcium and magnesium concentrations in drinking-water for health reasons should comply with the technical requirements to provide water suitable for distribution and should not compromise disinfection. Based on local circumstances and prevalent deficiencies, water suppliers and public health authorities may wish to further modify drinking-water composition in light of overall mineral nutrition.

Consumers should be informed of the mineral composition of their water, whether it is or is not modified. The contribution of drinking-water minerals to mineral nutrition should be considered where changes in supply are proposed or where less traditional sources, such as recycled water, seawater or brackish water, are processed and exploited for drinking-water. All of those techniques require that the water be stabilized prior to distribution, and addition of lime is a common and low-cost method.

Drinking-water may be a contributor of calcium and magnesium in the diet and could be important for those who are marginal for calcium and magnesium intake. Where

drinking-water supplies are supplemented with or replaced by demineralized water that requires conditioning, consideration should be given to adding calcium and magnesium salts to achieve concentrations similar to those that the population received from the original supply. As lime softening is commonly practised, addition of calcium and sometimes magnesium is a common practice for technical reasons and actually also adds calcium and often magnesium, which may be beneficial to dietary intake. Naturally soft water also often requires stabilization for corrosion control, and similar mineralization treatments should also be considered.

Although there is some evidence from epidemiological studies for a protective effect of magnesium or hardness on cardiovascular mortality, the evidence is being debated and does not prove causality. Further studies are being conducted. There are insufficient data to suggest either minimum or maximum concentrations of minerals at this time, and so no guideline values are proposed.

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